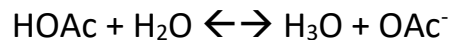


### Weak Acid/Base Practice Problem #1

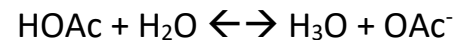
You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, H<sub>3</sub>O<sup>+</sup>, OAc<sup>-</sup>, and the pH if K<sub>a</sub> = 1.8x10<sup>-5</sup>.



Rxn	HOAc	↔	H <sub>3</sub> O	+	OAc <sup>-</sup>
I					
C					
E					
5%					
Answer					

### Weak Acid/Base Practice Problem #1

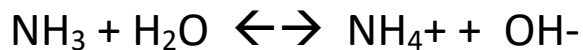
You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, H<sub>3</sub>O<sup>+</sup>, OAc<sup>-</sup>, and the pH if K<sub>a</sub> = 1.8x10<sup>-5</sup>.



Rxn	HOAc	↔	H <sub>3</sub> O	+	OAc <sup>-</sup>
I					
C					
E					
5%					
Answer					

### Weak Acid/Base Practice Problem #2

You have 0.010 M NH<sub>3</sub>. Calculate the pH. K<sub>b</sub> = 1.8 x 10<sup>-5</sup>

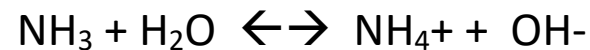


Rxn	NH <sub>3</sub>	↔	NH <sub>4</sub> <sup>+</sup>	+	OH <sup>-</sup>
I					
C					
E					
5%					
Answer					

**N-48**

### Weak Acid/Base Practice Problem #2

You have 0.010 M NH<sub>3</sub>. Calculate the pH. K<sub>b</sub> = 1.8 x 10<sup>-5</sup>



Rxn	NH <sub>3</sub>	↔	NH <sub>4</sub> <sup>+</sup>	+	OH <sup>-</sup>
I					
C					
E					
5%					
Answer					

**N-48**