Weak Acid/Base Practice Problem #1

You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, H_3O^+, OAc^-, and the pH if Ka = $1.8x10^{-5}$.

$$HOAc + H_2O \leftrightarrow H_3O + OAc^-$$

Rxn	HOAc ↔	H_3O +	OAc^-
I			
C			
E			
5%			
Answer			

Weak Acid/Base Practice Problem #2

You have 0.010 M NH_3. Calculate the pH. $K_b = 1.8 \times 10^{-5}$

$$NH_3 + H_2O \leftrightarrow NH_4+ + OH-$$

Rxn	NH_3 ↔	NH_4+ +	OH-
I			
C			
E			
5%			
Answer			

Weak Acid/Base Practice Problem #1

You have 1.00 M HOAc. Calc. the equilibrium concentrations of HOAc, H_3O^+, OAc^-, and the pH if Ka = $1.8x10^{-5}$.

$$HOAc + H_2O \leftrightarrow H_3O + OAc^-$$

Rxn	HOAc ↔	H_3O +	OAc^-
I			
C			
E			
5%			
Answer			

Weak Acid/Base Practice Problem #2

You have 0.010 M NH_3. Calculate the pH. $K_b = 1.8 \times 10^{-5}$

$$NH_3 + H_2O \leftrightarrow NH_4+ + OH-$$

Rxn	NH_3 ↔	NH_4+ +	OH-
I			
C			
E			
5%			
Answer			

N-48